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students and parents). 4 Question 2 2 Fe(s) +

3 2 O<sub>2</sub>(g) → Fe<sub>2</sub>O<sub>3</sub>(s) ΔH<sub>fD</sub> = -824 kJ mol<sup>-1</sup>

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$\text{CH}_3\text{CH}_2\text{COOH}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{CH}_3\text{CH}_2\text{COO}^-$   
 $(\text{aq}) + \text{H}_3\text{O}^+(\text{aq})$  Propanoic acid,  $\text{CH}_3\text{CH}_2\text{COOH}$ , is a carboxylic acid that reacts with water according to the equation above. At 25 °C the pH of a 50.0 mL sample of 0.20 M  $\text{CH}_3\text{CH}_2\text{COOH}$  is 2.86.

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Question 8 Answer the following questions about carbon monoxide,  $\text{CO}(\text{g})$ , and carbon dioxide,  $\text{CO}_2(\text{g})$ . Assume that both gases exhibit ideal behavior. (a) Draw the complete Lewis structure (electron-dot diagram) for the  $\text{CO}$  molecule and for the  $\text{CO}_2$  molecule. 1 point for each correct, complete Lewis structure

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