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of three beakers contains 25.0 mL  
of a 0.100 M solution of HCl, NH  
3, or NH 4 Cl, as shown above.

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student is assigned the task of  
determining the mass percent of

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silver in an alloy of copper and silver by dissolving a sample of the alloy in excess nitric acid and then precipitating the silver as AgCl. First the student prepares 50. mL of 6 M HNO<sub>3</sub> .

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resulting in a higher boiling point.  $\text{N}_2\text{H}_4$  is a polar molecule with London dispersion forces, dipole-dipole forces, and hydrogen bonding between molecules, whereas  $\text{C}_2\text{H}_6$  is nonpolar and only has London dispersion forces between

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molecules. It is 11% to the percent (c) The normal boiling point of  $N_2H_4$  is  $114.0^\circ C$ , whereas the normal boiling point of  $N_2O$  is  $-88.5^\circ C$ . Explain the difference in boiling points of the two molecules.

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THE NEXT PAGE. -7- 2. An 8.55  
mol sample of methanol, CH<sub>3</sub>OH  
, is placed in a 15.0 L evacuated  
rigid tank and heated to 327°C.

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thing you can do is remain calm and stay focused and methodical in your approach to each question.

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question #5 from the 2011 AP Chemistry exam. The question deals with Lewis dot-diagrams, Intermolecular Forces, reactions with water...

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6. In an experiment, all the air in  
a rigid 2.0 L flask is pumped out.  
Then some liquid ethanol is  
injected into the sealed flask,

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which is held at 35 ° C.

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we got word from the College  
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