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Review: Thermodynamics! AP
Chemistry: 6.1-6.5 Energy Diagrams,
Thermal Equilibrium, and Heat
Capacity Chapter 6 Thermochemistry
Review ~~Chapter 6 (Thermochemistry)~~
~~-Part 3 Chapter 6: Thermochemistry~~
Ch.6 Thermochemistry - 6.1 (The
Nature of Energy) Introduction to
chapter 6: thermochemistry AP
CHEMISTRY - Chapter 6:
Thermochemistry ~~Chapter 6 Lesson 1~~
~~Thermochemistry~~ Calorimetry
Concept, Examples and
Thermochemistry | How to Pass
Chemistry AP Chem CH7 Atomic
Structure and Periodicity Enthalpy of

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Reaction Chapter 16 (Spontaneity,
Entropy and Free Energy) - Part 2
Hess's Law and Heats of Formation

Section 11 - Thermo Chemistry (Part
2) Intro to Thermochemistry Tricks to
solve Thermochemistry problems
easily | Enthalpy of formation
combustion Chapter 7 (Atomic
Structure and Periodicity) - Part 1
Chapter 5 (Gases) - Part 1 Chapter 6
(Thermochemistry) - Part 2

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Chapter 6 Thermochemistry Energy
Flow and Chemical Change part 1

Chapter 6 thermochemistry 1 34

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6 Zumdahl Chapter 6 -

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162 CHAPTER 6:

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THERMOCHEMISTRY To convert the answer to joules, we write: 101.3 J

0.18 L atm $1 \text{ L atm} = - \quad \times = \quad \text{w}$

$- 18 \text{ J}$ 6.17 An expansion implies an increase in volume, therefore w must be $- 325 \text{ J}$ (see the defining equation for pressure-volume work.)

Chapter 6 Thermochemistry Test Answers

Chemistry Chapter 6 -

Thermochemistry. energy. radiant energy. thermal energy. chemical energy. the capacity to do work.

comes from the sun and is earth's primary energy source. the energy associated with random motion of atoms and molecules. energy stored within the bonds of chemical substances.

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Thermochemistry MULTIPLE CHOICE.
Choose the one alternative that best

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completes the statement or answers the question. ΔH $4\text{Al (s)} + 3\text{O}_2 \text{ (g)} \rightarrow 2\text{Al}_2\text{O}_3 \text{ (s)}$ A) exothermic, released B) exothermic, absorbed. AP Chemistry Practice Test, Ch. 6: Thermochemistry ... Start studying CHAPTER 6 THERMOCHEMISTRY TEST. Learn vocabulary, terms,

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MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question. 1) A chemical reaction that absorbs heat from the surroundings is said to be _____ and has a _____ ΔH at constant pressure. A) endothermic, positive

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Thermochemistry ...

164 CHAPTER 6:

THERMOCHEMISTRY Substituting into the above equation: $E = 483.6 \times 10^3 \text{ J} - (8.314 \text{ J/mol K})(398 \text{ K})(+1 \text{ mol})$ $E = 4.80 \times 10^5 \text{ J} = 4.80 \times 10^2 \text{ kJ}$ 6.28 We initially have 6 moles of gas (3 moles of chlorine and 3 moles of hydrogen). Since our product is 6

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MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question. 1) A chemical reaction that absorbs heat from the surroundings is said to be _____ and has a _____ ΔH at constant pressure. A) endothermic, positive.

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Answers

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reserved. 115 14. Two metals of equal
mass with different heat capacities are
subjected to the same amount of heat.
Which undergoes the smallest change

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in temperature? a) The metal with the higher heat capacity. b) The metal with the lower heat capacity.

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Ch.6 - Thermochemistry Ch.6.1: The
Nature of Energy Energy: An object ' s

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capacity to perform work or produce heat
Potential Energy: Energy due to position or composition (chemical bonds). Kinetic Energy: Energy due to the motion of the object
 $KE = \frac{1}{2}mv^2$
Law of Conservation of Energy: Energy can neither be created nor destroyed,

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19 - Atomic Structure & Periodicity,
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Materials. ... Most of the content in
these relates strongly to this chapter.
Anything that is not included in what
we studied you would not be

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responsible for.

Chapter 6 - Thermochemistry - Mrs.
Duffey - FHN

Chapter 6 Thermochemistry flow is
determined into or out of the
surroundings. Because $E_{univ} = 0$ by
the first law of thermodynamics,

$E_{sys} = -E_{surr}$; what happens to
the surroundings is the exact opposite
of what happens to the system.

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